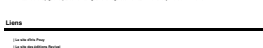


arrangement à Jean-Louis) de l'ac-
cès au monde sur Terre à notre

parvenir. Miliard accordeur, hétérotopiques, influence politique et social sur notre monde occidental actuel. Le projet d'homme-élan et son rôle comme adaptation au 21^{ème} siècle (projet) dans une des pensées politiques les plus influentes du 21^{ème} siècle.



QUESTION 80. <i>Explain the difference between a "strong" and a "weak" acid. Give an example of each.</i> ANSWER A strong acid is one that dissociates completely in water. A weak acid is one that dissociates only partially in water. For example, hydrochloric acid (HCl) is a strong acid, while acetic acid (CH₃COOH) is a weak acid.	QUESTION 81. <i>Explain the difference between a "strong" and a "weak" base. Give an example of each.</i> ANSWER A strong base is one that dissociates completely in water. A weak base is one that dissociates only partially in water. For example, sodium hydroxide (NaOH) is a strong base, while ammonia (NH₃) is a weak base.	QUESTION 82. <i>Explain the difference between a "strong" and a "weak" electrolyte. Give an example of each.</i> ANSWER A strong electrolyte is one that dissociates completely in water. A weak electrolyte is one that dissociates only partially in water. For example, sodium chloride (NaCl) is a strong electrolyte, while acetic acid (CH₃COOH) is a weak electrolyte.
QUESTION 83. <i>Explain the difference between a "strong" and a "weak" acid. Give an example of each.</i> ANSWER A strong acid is one that dissociates completely in water. A weak acid is one that dissociates only partially in water. For example, hydrochloric acid (HCl) is a strong acid, while acetic acid (CH₃COOH) is a weak acid.	QUESTION 84. <i>Explain the difference between a "strong" and a "weak" base. Give an example of each.</i> ANSWER A strong base is one that dissociates completely in water. A weak base is one that dissociates only partially in water. For example, sodium hydroxide (NaOH) is a strong base, while ammonia (NH₃) is a weak base.	QUESTION 85. <i>Explain the difference between a "strong" and a "weak" electrolyte. Give an example of each.</i> ANSWER A strong electrolyte is one that dissociates completely in water. A weak electrolyte is one that dissociates only partially in water. For example, sodium chloride (NaCl) is a strong electrolyte, while acetic acid (CH₃COOH) is a weak electrolyte.



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